Maharashtra Educational Society's

H. K. College of Pharmacy

Oshiwara,Jogeshwari(W)- 102

Question Bank

Pharmaceutical Analysis I

Sem I (Syllabus R-2019)

Note: This is just a sample question bank to get an idea about the kind of questions asked in final online exam based on MCQs. Final exam MCQs can have relevance but questions would be totally different than sample question bank.

	Option a	**Option b**	**Option c**	**Option d**
Q.1	Impurities in pharmaceutical preparation are may be due to following sources:			
	Raw material	Manufacturing process	Chemical instability	All of the above
Q 2	Primary standard used for standardisation of NaOH			
	Lead acetate	Potassium dichromate	Potassium hydrogen phthalate	All of the above
Q 3	Normality is			
	Moles of solute per litre of solution	Gram per equivalent weight of solute per litre of the solution	Parts per million	Grams per ml
Q 4	Standardization of sodium thiosulphate is a titration of			
	Acid -Base type	Iodometry	Nonaqueous	Iodimetry
Q 5	Systematic errors are of the following type except			
	Personal error	Method error	Random error	Reagent error

	Option a	**Option b**	**Option c**	**Option d**
Q.6	% w/w express the			
	Number of grams of solute in 1000 gm of product	Number of grams of solute in 100 gm of product	Number of ml of solute in 100 ml of product	Number of grams of solute in 100 ml of product

Q 7	The endpoint of an EDTA titration is usually found by using a _____ indicator.			
	Metallochromic	Redox	Acid-Base	All
Q 8	Titrations based on the use of silver nitrate are called _____titration.			
	Complexometric	Argentometric	Amperometric	Conductometric
Q 9	Adsorption indicators are used in _____			
	Fajan's method	Mohr's method	Volhard's method	All
Q 10	EDTA forms ____ complex with the metal ion. Buffer (NH_3-NH_4Cl) added maintains the pH around_____			
	1:2, 10.5	1:1, 10.5	2:1, 2.5	1:1, 2.5
Q 11	Given: Ksp AgCl= 1.2 X 10^{-10} and Ksp AgI= 1.7 X 10^{-16}. In a solution of Cl^- and Br^-, when Ag^+ solution is gradually added, which halide will precipitate first?			
	AgCl	AgI	Both	There will be no precipitation
Q 12	Assay of Sulphacetamide sodium is an example of _____			
	Dizotization reaction	Nitrite titration	both a and b	None of the above
Q 13	The diffusion current in the polarography depends on all of the following, EXCEPT			
	Capillary diameter	Life time of mercury drop	Temperature	Charge of the electrolyte
Q 14	Starch Iodide paper used in nitrite titration is _____			
	used to check pH of titrant	used to check pH of titrand	used as an external indicator	None of the above
Q 15	EDTA has _____ binding sites and therefore it is also called as multidentate ligand.			
	3	4	5	6
Q 16	Protophilic solvents are			
	NH_3	CCl_3	HF	Acetic acid
Q 17	Nonaqueous titration are based on			
	Arrhenius concept	Lowry-Bronsted concept	Lewis concept	None of the above

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Q 18	Basicity of organic acids can be determined by			
	complexometry	potentiometry	conductometry	None of the above
Q 19	Equivalent conductance is…… related with concentration			
	Directly	Inversely	Logarithmically	All of the above
Q 20	Mode of a set of data is the number with			
	Lowest frequency	Highest frequency	Both a and b	None of the above
Q 21	A _____ is the electrode whose potential is known and remains constant			
	reference electrode	indicator electrode	pH electrode	Graphite electrode
Q 22	The electrode potentials are calculated by			
	Ilkovi equation	Nernst equation	Stokes equation	Ohm's law
Q 23	As temperature increases electrolytic conduction			
	increases	decreases	remains unaffected	none of the above
Q 24	Find the oxidation state of I in H4IO6–			
	+7	+5	+1	-1
Q 25	The indicator used in iodometric titrations is			
	Phenolphthalein	KI	litmus	starch
Q 26	Which of the following is not a property required of the substance chosen for use as a precipitate in a gravimetric analysis?			
	Low solubility.	Stable when heated to 110°C.	Able to be stored for an extended time without deterioration.	Has known formula.